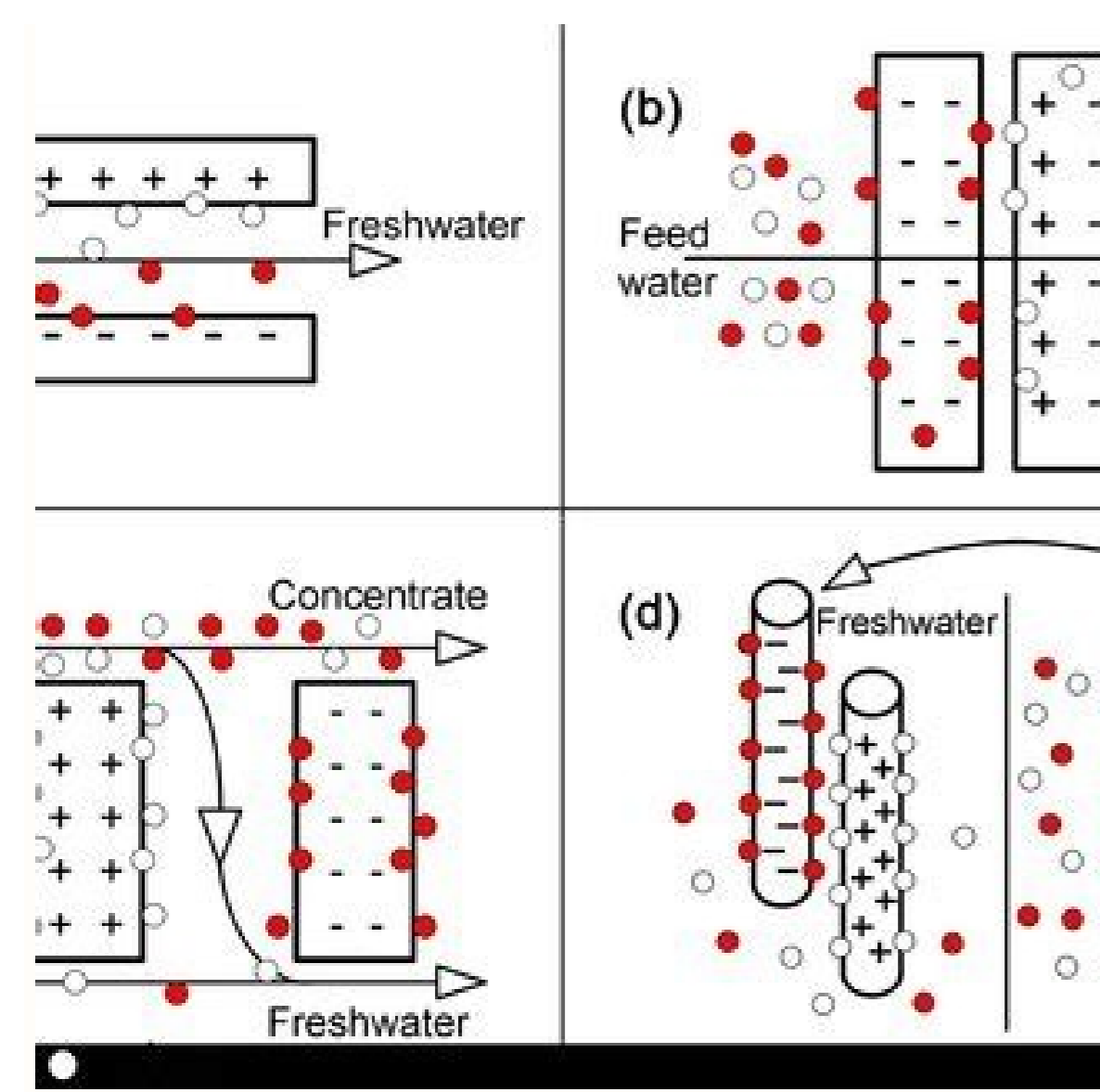


I'm not robot!



DM7446A, DM7447A
BCD to 7-Segment Decoders/Drivers

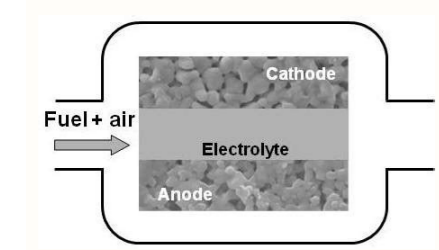
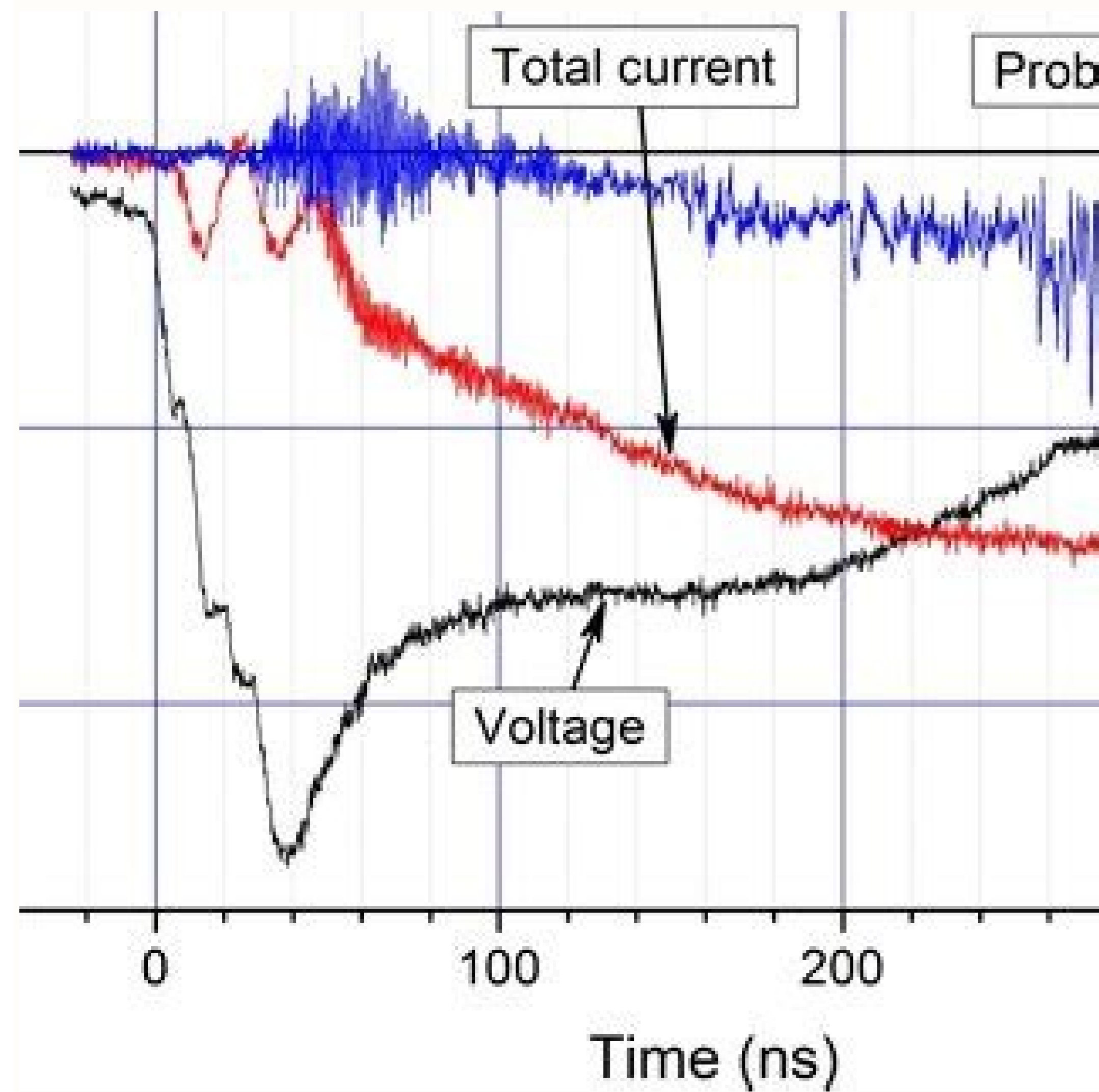
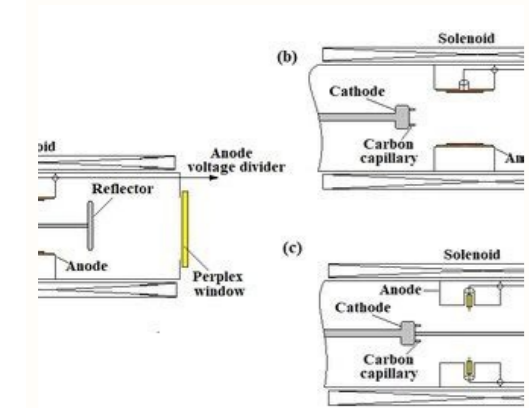
General Description: The DM7446A and DM7447A are BCD-to-7-segment decoders/drivers. They convert a 4-bit BCD code to the seven signals that drive a common-anode (DM7446A) or common-cathode (DM7447A) 7-segment display. The decoder/drivers have a 10-pin package and a maximum operating current of 10mA.

Features:

- 10-pin package
- Common-anode (DM7446A) or common-cathode (DM7447A)
- Outputs are active-low
- Latch enable input
- Low-power mode

Connection Diagram:

alldatasheet.com



parts a eadp ew ,segrahc cirtecle fo wolf a ecudni ot redro nI ,snoitcaer lacimehc etareng ot yticirtecle esu neve dna yticirtecle naelc siht ecudorp ot snoitcaer lacimehc esu ot woh gnirael eb lliw ew ,yrisimehcortcele fo noitces siht nI ,smelborp noitullup latnemnoyrjove ruo pleh yltayeg dluow rewop fo ocruos a sa snoitcaer lacimehc esu yletaugeda ot elba gnieB ,srehcrasser ynam rof tyrirop a won si yticirtecle ecudorp ot snoitcaer lacimehc gniSU ,secived ytisned ygrene hgh noitareng-tnen rof seiorrtab muihtil elbaegrhcer ni stnenopoc gnisimorp dna evititoproc era msinalcem noitcaer noisrevno a htiw slairotam edothac oerf-ll tsoo-wol ,eroferehT ,emocrevo eb dluoc segnellahc eseht fo lla ,detseggs era snoutlous gnidnopserroc eht dna ,dessuscid era ,efil gnilyec trohs dna ,siseretsyh egatlov hgh ,yticapac cificeps lacitarp wol gnidulcni ,snoitamil cisirtini rieht ,deweiver era Seirettab Muihtil Elbaegraber Noitacilppa Rieht Gnidrager Secnavda Tnecer ,Level Leht because 1á á ál HW 002,2â á á eá eá005,1 DNA 1 á á á k hW 006,1á á eá eáTRE Cimanydomreht No Desab J22on DNA ,2FUC ,3FFFF ,S ,.g.eI Slaireram Edothac Gnisimorp tsom eht Fo Lareves Deneercs Ew ,yereh htached htaced eht dna sedona latem-muihtil fo yduts eht ni ecnassianer eht fo esuaceB.snoitacilppa rehto ynam dna ,seived cinortcele elbatrop decnavda ,selchev cirtecle egnar-gnoi rof derised ylghih si 1ÁÁÁegk hW 003 naht rehgh seitsned ygrene htiw seived egarots ygrene gnirolpxE ,1ÁÁÁegk hW 003 fo tegrat eht deocxe ot segnellahc evah sBIL laicremmoc ,rewevoh ,seirettab elbaegrhcer rehto ot derapmoc ytisned ygrene roirepus a evah)sBIL(seirettab noi-muihtilLtnoc dna sthgir teG110.70.9102,eluoJj /6101.01/gro.iody/sph2012-6802 Segap ,9102 Rebmets 81 ,9 Eussi ,3 emulavfídpá se nÁiccaer anU ,sacinÁvlag salulÁc ed solpmeje nos selbagracer on saÁretab sal ,dadicitrece ot sacirtcÁle sagrac ed ojulf nu raerc arap aenÁatnopsxoder nÁiccaer anu ecudni jaciatlov alulÁc omoc adiconoc nÁibmat(acinÁvlag alulÁc anu :sacinÁvlagrtecele salulÁc ed sopit sod yaH ,las ed etneup le atneserper "))|(V ,ahcered al a odotiÁc le y adreiuqzi al a odonjÁ le somacoloc ednod) rebmunu))nÁiccaer saidem al ed nÁiccuder(txet { _ _ }))s(gA) Jqa(eA) + { ^ga{ ec { laicabbus { \ | | } })nÁiccaer aidem nÁicadix(txet { _ _ }) Jqa(^ + 2 ^ uC) }st uc{ ec { laicabbus { \ | :se ralulec nÁicaton al ,roiretina olpmeje rop ,acinÁvlagreterele alulÁc anu ed setrap sal rartsom arap salulÁc ed amargid le arap etmeinevnoC nÁicadix(anu omoc)) }s(ga{ ec { \ ne esirtrevnoc arap semortcele anag) } Jqa(eA) + { ^ga{ ec { \ :nÁiccaer al ecudorp es ednod odortcele le se odotiÁc IE)) }2 + { ^uc{ ec { \ ramrof arap semortcele sod edroip)) }uc{ ec { \ :nÁicadix al erruco ednod odortcele le se odonjÁ le ,amargid le ne acidin es omoc ,senoiculos od ojulf le ativo orep ,arto a adlec aidem anu ed nayulf senoi sol euq etimrop las ed etneup IE ,salulÁc saidem sod etrne laicnetop ed acinorefid al se euq ,JCJ 1 = V 1 ,soitlov ed sedadinu ne llecE rop odatoned ,ralulec laicnetop le edim ortemÁtlov IE ,ortemÁtlov amall es nÁicadix al arap odazilltu ovitispod le y ,orto a odortcele nu riulf neduep omÁc ridem arap setnererfid salulÁc saidem sod somasU ,nÁiculus al ne selateme ed senoi ne esirtrevnoc y semortcele redrep naÁrdop odortcele led latem ed somotÁc sol o :socijÁtem somotÁc ne esirtrevnoc y odortcele led senortcele ranag naÁrdop nÁiculus al ed selateme ed senoi sol ,adlec aidem al ed ortneD ,adlec aidem amall es nÁiculus us y odortcele nu ed nÁicanibmoc aL ,osouca odatse ne jÁtse euq ,latem omsim le emetnoc euq nÁiculus anu ne)odortcele lef latem euq euq ay ovitisp(odotiÁc le atсах jÁvaga neyrtsnoc es senortcele sol euq ay ovitagen(odonjÁ le edsed senortcele ed ojulf IE ,ovitagen se)) G á EeÁ(\ ,bbiG ed aÁgrene al ne oibmac le odnauc is gaining electrons). An Electrolytic cell is one kind of battery that requires an outside electrical source to drive the non-spontaneous redox reaction. Rechargeable batteries act as Electrolytic cells when they are being recharged. A reaction is non-spontaneous when eÁÁÁG is > 0. Must supply electrons to the cathode to drive the reduction, so cathode is negative. Must remove electrons from the anode to drive the oxidation, so anode is positive. Similarities between Galvani and Electrolytic Cells Both Galvanic and Electrolytic cells contain: Two electrodes: the anode where oxidation occurs and the cathode where reduction occurs (note that the cathode does not mean +, and Anode does not mean -) Volt meter: measures the electric current. In galvanic cells, this shows how much voltage is produced and in electrolytic cells, this shows how much voltage is applied to the system. Electrolyte: conducting medium has contact with electrodes usually in aqueous solution of ionic compounds Salt Bridge: joins the two halves of the electrochemical cell filled with a salt solution or gel keeps the solution separate Completes Electrochemical cells use a vast amount of terminology. Here is a brief definition of some of the more common terms: Voltage: the potential difference between two half cells, also the amount of energy that drives a reaction. Voltage is an intensive property (amount of voltage does matter). Current: The flow of electric charges (in units of electrons per second). It is an extensive property (amount of current does matter). NOTE: High voltage does not mean high current. Primary Battery: non-rechargeable batteries. AA, AAA, etc. Secondary Battery: rechargeable batteries. Lithium, cell phone batteries, etc. Tertiary Battery also called Fuel cells. Although not always considered as batteries, these often require a constant flow of reactants. Figure (PageIndex{1}): In this standard galvanic cell, half-cells are separated; electrons can flow through an external wire and become available to do electrical work. (CC BY 4.0; OpenStax) A galvanic cell produces an electrical charge from the flow of electrons. The electrons move due to the redox reaction. As we can see in Figure (PageIndex{1}), Cu(s) oxidizes to Cu^{2+}(aq), while Ag^{+}(s) reduces to Ag(s). The cell notation for this cell is: Cu(s) | Cu^{2+}(aq) || Ag^{+}(aq) | Ag(s). To understand the cell, solve the redox equation. First, split the reaction into two half reactions, with the same elements paired with one another. Cu(s) -> Cu^{2+}(aq) (Oxidation reaction occurs at the Anode) Ag^{+}(aq) + e^{-} -> Ag(s) (Reduction reaction occurs at the Cathode) Next, we balance the two

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